

$$\left(\frac{\partial u}{\partial v}\right)_T = T \left(\frac{\partial p}{\partial T}\right)_v - p$$

$$\left(\frac{\partial u}{\partial v}\right)_T = T \left(\frac{\partial p}{\partial T}\right)_v - \frac{RT}{v-b} + \frac{a}{v^2}$$

Van der Waals gas:

$$\left(p + \frac{a}{v^2}\right)(v-b) = RT$$

$$p = \frac{RT}{v-b} - \frac{a}{v^2}$$

$$= \frac{RT}{v-b} - \frac{RT}{v-b} + \frac{a}{v^2} = \frac{a}{v^2}$$

~~$$\left(\frac{\partial u}{\partial v}\right)_T \left(\frac{\partial v}{\partial T}\right)_u \left(\frac{\partial T}{\partial u}\right)_v = -1$$~~

~~$$-\left(\frac{\partial u}{\partial v}\right)_T \left(\frac{\partial v}{\partial T}\right)_u = \left(\frac{\partial u}{\partial T}\right)_v = C_v$$~~

To show: $C_v = C_v(T)$

$$C_v = \left(\frac{dq}{dT}\right)_v$$

$$\left(\frac{\partial p}{\partial T}\right)_v = \left(\frac{\partial}{\partial T} \left[\frac{RT}{v-b} - \frac{a}{v^2} \right]\right)_v = \frac{R}{v-b}$$

$$\left(\frac{\partial u}{\partial v}\right)_T = \frac{a}{v^2}$$

$$\partial u = \frac{a}{v^2} \partial T$$

$$u = \int \frac{1}{v} a + u_0(T) = \frac{a}{v} + u_0(T)$$

$$C_v = \left(\frac{\partial u}{\partial T}\right)_v = \left(\frac{\partial u_0(T)}{\partial T}\right)_v = C_v(T)$$